

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization International Bureau



(43) International Publication Date
29 December 2004 (29.12.2004)

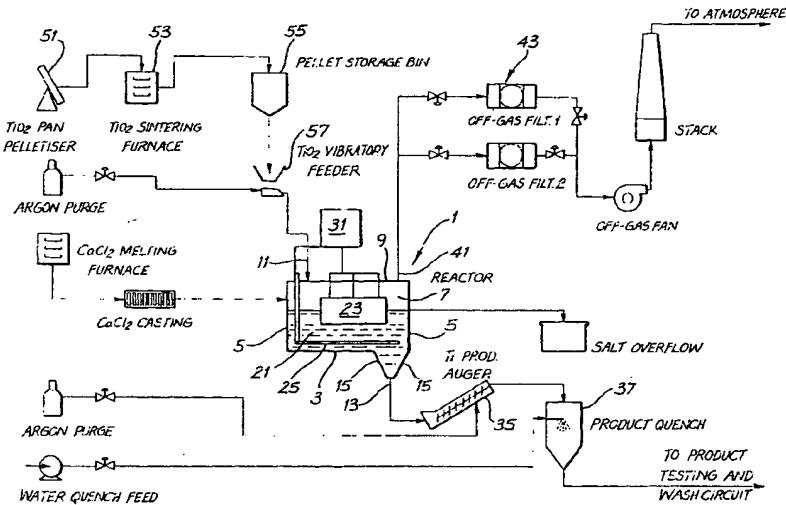
PCT

(10) International Publication Number
WO 2004/113593 A1

- (51) International Patent Classification⁷: C25C 5/00, RATCHEV, Ivan [AU/AU]; 2 Young Street, Georgetown, New South Wales 2298 (AU).
- C22B 9/14, 34/12
- (21) International Application Number: PCT/AU2004/000809 (74) Agent: GRIFFITH HACK; 509 St Kilda Road, Melbourne, Victoria 3004 (AU).
- (22) International Filing Date: 21 June 2004 (21.06.2004) (81) Designated States (unless otherwise indicated, for every kind of national protection available): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NA, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SY, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.
- (25) Filing Language: English
- (26) Publication Language: English
- (30) Priority Data: 2003903150 20 June 2003 (20.06.2003) AU
- (71) Applicant (for all designated States except US): BHP BILLITON INNOVATION PTY LTD [AU/AU]; BHP Billiton Centre, 180 Lonsdale Street, Melbourne, Victoria 3000 (AU).
- (72) Inventors; and
- (75) Inventors/Applicants (for US only): SHOOK, Andrew, Arthur [AU/AU]; 4 Orlando Close, Eleebana, New South Wales 2282 (AU). RIGBY, Gregory, David [AU/AU]; 107 Tirriki Street, Charles Town, New South Wales 2290 (AU).
- (84) Designated States (unless otherwise indicated, for every kind of regional protection available): ARIPO (BW, GH, GM, KE, LS, MW, MZ, NA, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PL, PT, RO, SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

[Continued on next page]

(54) Title: ELECTROCHEMICAL REDUCTION OF METAL OXIDES



WO 2004/113593 A1

(57) Abstract: An electrolytic cell for electrochemically reducing metal oxide powders and/or pellets is disclosed. The cell includes a cathode (25) in the form of a plate that has an upper surface for supporting metal oxide powders and/or pellets. The plate is horizontally disposed or slightly inclined and has a forward end and a rearward end and is immersed in an electrolyte bath. The plate is supported for movement so as to cause metal oxide powders and/or pellets on the upper surface of the plate to move toward a forward end of the plate. The cell also includes a means for causing metal oxide powders and/or pellets to move over the upper surface of the plate toward the forward end of the cathode while in contact with molten electrolyte whereby electrochemical reduction of the metal oxide to metal can occur. A method of continuously or semi-continuously reducing metal oxide powders and/or pellets in the cell is also disclosed.



Published:

— *with international search report*

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

- 1 -

ELECTROCHEMICAL REDUCTION OF METAL OXIDES

The present invention relates to electrochemical reduction of metal oxides.

5

The present invention relates particularly to continuous and semi-continuous electrochemical reduction of metal oxides in the form of pellets to produce metal having a low oxygen concentration, typically no more than 10 0.2% by weight.

The present invention was made during the course of an on-going research project on electrochemical reduction of metal oxides being carried out by the 15 applicant. The research project has focussed on the reduction of titania (TiO_2).

During the course of the research project the applicant carried out experimental work on the reduction 20 of titania using electrolytic cells that included a pool of molten $CaCl_2$ -based electrolyte, an anode formed from graphite, and a range of cathodes.

The $CaCl_2$ -based electrolyte was a commercially 25 available source of $CaCl_2$, namely calcium chloride dihydrate, that decomposed on heating and produced a very small amount of CaO .

The applicant operated the electrolytic cells at 30 a potential above the decomposition potential of CaO and below the decomposition potential of $CaCl_2$.

The applicant found that at these potentials the cell could electrochemically reduce titania to titanium 35 with low concentrations of oxygen, ie concentrations less than 0.2 wt.%.

- 2 -

The applicant does not have a complete understanding of the electrolytic cell mechanism at this stage.

5 Nevertheless, whilst not wishing to be bound by the comments in this paragraph and the following paragraphs, the applicant offers the following comments by way of an outline of a possible cell mechanism.

10 The experimental work carried out by the applicant produced evidence of Ca metal dissolved in the electrolyte. The applicant believes that the Ca metal was the result of electrodeposition of Ca^{++} cations as Ca metal on the cathode.

15 As is indicated above, the experimental work was carried out using a CaCl_2 -based electrolyte at a cell potential below the decomposition potential of CaCl_2 . The applicant believes that the initial deposition of Ca metal 20 on the cathode was due to the presence of Ca^{++} cations and O^{-} anions derived from CaO in the electrolyte. The decomposition potential of CaO is less than the decomposition potential of CaCl_2 . In this cell mechanism the cell operation is dependent on decomposition of CaO , 25 with Ca^{++} cations migrating to the cathode and depositing as Ca metal and O^{-} anions migrating to the anode and forming CO and/or CO_2 (in a situation in which the anode is a graphite anode) and releasing electrons that facilitate electrolytic deposition of Ca metal on the 30 cathode.

The applicant believes that the Ca metal that deposits on the cathode directly or indirectly (via dissolution of Ca metal in the electrolyte) participates 35 in chemical reduction of titania resulting in the release of O^{-} anions from the titania.

- 3 -

The applicant also believes that the O⁻ anions, once extracted from the titania, migrate to the anode and react with anode carbon and produce CO and/or CO₂ (and in some instances CaO) and release electrons that facilitate
5 electrolytic deposition of Ca metal on the cathode.

The applicant operated the electrolytic cells on a batch basis with titania in the form of pellets and larger solid blocks in the early part of the work and
10 titania powder in the later part of the work. The applicant also operated the electrolytic cells on a batch basis with other metal oxides.

Whilst the research work established that it is
15 possible to electrochemically reduce titania (and other metal oxides) to metals having low concentrations of oxygen in such electrolytic cells, the applicant has realised that there are significant practical difficulties
20 operating such electrolytic cells commercially on a batch basis.

In the course of considering the results of the research work and possible commercialisation of the technology, the applicant realised that commercial
25 production could be achieved by operating an electrolytic cell on a continuous or semi-continuous basis with metal oxide powders and pellets being transported through the cell in a controlled manner and being discharged in a reduced form from the cell.

30 International application PCT/AU03/001657 lodged on 12 December 2003 in the name of the applicant describes this invention in broad terms as a process for electrochemically reducing a metal oxide, such as titania,
35 in a solid state in an electrolytic cell that includes a bath of molten electrolyte, a cathode, and an anode, which process includes the steps of: (a) applying a cell

- 4 -

potential across the anode and the cathode that is capable of electrochemically reducing metal oxide supplied to the molten electrolyte bath, (b) continuously or semi-continuously feeding the metal oxide in powder and/or pellet form into the molten electrolyte bath, (c) transporting the powders and/or pellets along a path within the molten electrolyte bath and reducing the metal oxide to metal as the metal oxide powders and/or pellets move along the path, and (d) continuously or semi-continuously removing reduced metal oxide from the molten electrolyte bath.

The International application defines the term "powder and/or pellet form" as meaning particles having a particle size of 3.5 mm or less. The upper end of this particle size range covers particles that are usually described as pellets. The terms "powder" and "pellets" as used herein are not intended to limit the scope of patent protection to a particular procedure for producing the particles.

The term "semi-continuously" is understood in the International application and herein to mean that the process includes: (a) periods during which metal oxide powders and pellets are supplied to the cell and periods during which there is no such supply of metal oxide powders and pellets to the cell, and (b) periods during which metal is removed from the cell and periods during which there is no such removal of metal from the cell.

The overall intention of the use of the terms "continuously" and "semi-continuously" in the International application and herein is to describe cell operation other than on a batch basis.

In this context, the term "batch" is understood in the International application and herein to include

- 5 -

situations in which metal oxide is continuously supplied to a cell and reduced metal builds up in the cell until the end of a cell cycle, such as disclosed in International application WO 01/62996 in the name of The
5 Secretary of State for Defence.

The disclosure in the International application is incorporated herein by cross reference.

10 The applicant has carried out further research into commercial production by operating an electrolytic cell on a continuous or semi-continuous basis and has realised that the cell should include a cell cathode in the form of a member, such as a plate, having an upper
15 surface for supporting metal oxide particles in powder and/or pellet form that is horizontally disposed or slightly inclined (upwardly or downwardly) and has a forward end and a rearward end and is immersed in the electrolyte bath and is supported for movement, preferably
20 in forward and rearward directions, so as to cause metal oxide powders and/or pellets to move toward the forward end of the cathode.

With this arrangement, in use, metal oxide
25 powders and/or pellets are supplied onto the upper surface of the cathode, preferably near the rearward end thereof, and are moved forward by the movement of the cathode and fall off the upper surface at the forward end of the cathode and ultimately are removed from the cell. The
30 metal oxides are reduced as the metal oxides powders and/or pellets move over the upper surface.

The term "powders and/or pellets" is understood herein to mean particles that are less than 5mm in major
35 dimension.

Accordingly, the present invention provides a

- 6 -

process for electrochemically reducing metal oxide powders and/or pellets, such as titania powders and/or pellets, in an electrolytic cell that includes a bath of molten electrolyte, a cathode, and an anode, the cathode being in 5 the form of a member, such as a plate, having an upper surface for supporting metal oxide powders and/or pellets that is horizontally disposed or slightly inclined and has a forward end and a rearward end and is immersed in the electrolyte bath and is supported for movement so as to cause metal oxide powders and/or pellets on the upper 10 surface of the cathode to move toward the forward end of the member, which process includes the steps of: (a) applying a cell potential across the anode and the cathode that is capable of electrochemically reducing metal oxide supplied to the molten electrolyte bath, (b) continuously 15 or semi-continuously feeding metal oxide powders and/or pellets into the molten electrolyte bath so that the powders and/or pellets deposit on an upper surface of the cathode, (c) causing metal oxide powders and/or pellets to move over the upper surface of the cathode toward the forward end of the cathode while in contact with molten electrolyte whereby electrochemical reduction of the metal oxide to metal occurs as the powders and/or pellets move toward the forward end, and (d) continuously or semi- 20 continuously removing at least partially electrochemically reduced metal oxide powders and/or pellets from the molten electrolyte bath.

30 Preferably step (b) includes feeding the metal oxide powders and/or pellets into the molten electrolyte bath so that the powders and/or pellets form a layer that is one or two particles deep on the upper surface of the cathode.

35 The metal oxide powders and/or pellets may be deposited on the upper surface of the cathode in a pile of pellets and may be shaken out into one or two particle

- 7 -

deep layer as the cathode moves the powders and/or pellets towards the forward end of the cathode.

Preferably step (c) includes causing metal oxide
5 pellets to move on the upper surface of the cathode toward
the forward end of the cathode as a layer of powders
and/or pellets that is one or two particles deep.

The layer may be produced by forming the cathode
10 appropriately. For example, the cathode may be formed
with an upstanding lip at the forward end that causes
powders and/or pellets to build-up behind the lip.
Alternatively, or in addition, the cathode may be formed
15 with a series of transversely extending grooves that
promote close packing of the powders and/or pellets.

Preferably step (c) includes selectively moving
the cathode so as to cause metal oxide powders and/or
pellets on the upper surface of the cathode to move toward
20 the forward end of the cathode.

There is a wide range of options for moving the
cathode to cause forward movement of powders and/or
pellets on the upper surface of the cathode. The applicant
25 has found that it is preferable to move the cathode in
forward and rearward directions. The applicant has found
that one option that can achieve controlled forward
movement of powders and/or pellets includes moving the
cathode in a repeated sequence that comprises a short
30 period of oscillating motion in the forward and rearward
directions and a short rest period. The applicant has
found that this sequence can cause powders and/or pellets
on the upper surface of the cathode to move over the upper
surface in a controlled series of short steps from the
35 rearward end to the forward end of the cell. The
applicant has also found that controlled forward movement
of powders and/or pellets may include components of

- 8 -

rearward and forward movement of the controlled forward movement of powders and/or pellets, with a net forward movement.

5 Moreover, the present invention is not confined to operating a cell under constant operating conditions and extends to situations in which the operating parameters, such as the cathode movement, are varied during the operating campaign of the cell.

10 Preferably step (c) includes moving the cathode so as to cause powders and/or pellets across the width of the cathode to move at the same rate so that the powders and/or pellets have substantially the same residence time
15 within the bath.

20 Preferably the process electrochemically reduces the metal oxide to metal having a concentration of oxygen that is no more than 0.5% by weight.

25 More preferably the concentration of oxygen is no more than 0.2% by weight.

30 The process may be a single or multiple stage process involving one or more than one electrolytic cell.

In the case of a multiple stage process involving more than one electrolytic cell, the process may include successively passing reduced and partially reduced metal oxides from a first electrolytic cell through one or more than one downstream electrolytic cell and continuing reduction of the metal oxides in these cells.

35 In a situation in which the cathode is in the form of a plate, another option for a multiple stage process includes successively passing reduced and partially reduced metal oxide particles from one cathode

- 9 -

plate to another cathode plate or a succession of cathode plates within one electrolytic cell.

Another option for a multiple stage process
5 includes recirculating reduced and partially reduced metal oxide particles through the same electrolytic cell.

Preferably the process includes washing powders and/or pellets that are removed from the cell to separate 10 electrolyte that is carried from the cell with the powders and/or pellets.

The process inevitably results in a loss of electrolyte from the cell and, therefore make-up 15 electrolyte will be required for the cell.

The make-up electrolyte may be obtained by recovering electrolyte that is washed from the powders and/or pellets and recycling the electrolyte to the cell.

20 Alternatively, or in addition, the process may include supplying fresh make-up electrolyte to the cell.

Preferably the process includes maintaining the 25 cell temperature below the vaporisation and/or decomposition temperatures of the electrolyte.

Preferably the process includes applying a cell potential above a decomposition potential of at least one 30 constituent of the electrolyte so that there are cations of a metal other than that of the cathode metal oxide in the electrolyte.

In a situation in which the metal oxide is 35 titania it is preferred that the electrolyte be a CaCl₂-based electrolyte that includes CaO as one of the constituents.

- 10 -

In such a situation it is preferred that the process includes maintaining the cell potential above the decomposition potential for CaO.

5

Preferably the particle size of the powders and/or pellets is in the range of 0.5-4 mm.

More preferably the particle size of the pellets
10 is in the range of 1-2 mm.

According to the present invention there is also provided an electrolytic cell for electrochemically reducing metal oxide powders and/or pellets, which
15 electrolytic cell includes (a) a bath of a molten electrolyte, (b) a cathode in the form of a member, such as a plate, having an upper surface for supporting metal oxide powders and/or pellets that is horizontally disposed or slightly inclined and has a forward end and a rearward
20 end and is immersed in the electrolyte bath and is supported for movement so as to cause metal oxide powders and/or pellets on the upper surface of the cathode to move toward the forward end of the cathode, (c) an anode, (d) a means for applying a potential across the anode and the
25 cathode, (e) a means for supplying metal oxide powders and/or pellets to the electrolyte bath so that the metal oxide powders and/or pellets can deposit onto an upper surface of the cathode, (f) a means for causing metal oxide powders and/or pellets to move over the upper
30 surface of the cathode toward the forward end of the cathode while in contact with molten electrolyte whereby electrochemical reduction of the metal oxide to metal can occur as the powders and/or pellets move toward the forward end, and (g) a means for removing at least
35 partially electrochemically reduced metal oxides from the electrolyte bath.

- 11 -

Preferably the cathode is a plate.

Preferably the means for causing metal oxide powders and/or pellets to move over the upper surface of
5 the cathode includes a means for moving the cathode so as to cause movement of metal oxide powders and/or pellets.

10 Preferably the means for causing metal oxide powders and/or pellets to move over the upper surface of the cathode includes a means for moving the cathode in forward and rearward directions.

15 Preferably the cathode is formed to cause metal oxide powders and/or pellets to move on the upper surface of the cathode toward the forward end of the cathode as a layer that is one or two particles deep.

20 For example, the cathode may be formed with an upstanding lip at the forward end that causes pellets to build-up behind the lip. Alternatively, or in addition, the upper surface of the cathode may be formed with a series of transversely extending grooves that promote close packing of the pellets.

25

Preferably the means for applying an electrical potential across the anode and the cathode includes an electrical circuit in which a power source is connected to a forward end of the cathode. The applicant has found
30 that this arrangement results in substantial reduction of titania powders and/or pellets within a short distance from the forward end of the cell.

35 Preferably the anode extends downwardly into the electrolyte bath and is positioned a predetermined distance above the upper surface of the cathode.

- 12 -

In a situation in which the anode is a consumable anode, for example formed from graphite, preferably the cell includes a means for moving the anode downwardly into the electrolyte bath as the anode is consumed to maintain 5 the predetermined distance between the anode and the cathode.

More preferably the anode is in the form of one or more graphite blocks extending into the cell.

10

Preferably the cell includes a means for treating gases released from the cell.

15

The gas treatment means may include a means for removing any one or more of carbon monoxide, carbon dioxide, chlorine-containing gases such as phosgene from the gases.

20

The gas treatment means may also include a means for combusting carbon monoxide gas in the gases.

25

In a situation in which the metal oxide is titania it is preferred that the electrolyte be a CaCl_2 -based electrolyte that includes CaO as one of the constituents.

Preferably the particle size of the powders and/or pellets is in the range of 0.5-4 mm.

30

More preferably the particle size of the powders and/or pellets is in the range of 1-2 mm.

35

The present invention is described further by way of example with reference to the accompanying drawing which is a schematic diagram that illustrates one embodiment of an electrochemical process and an electrolytic cell in accordance with the present

- 13 -

invention.

The following description is in the context of electrochemically reducing titania pellets to titanium metal having an oxygen concentration of less than 0.3 wt.%. However, it is noted that the present invention is not confined to this metal oxide and extends to other metal oxides.

10 The electrolytic cell 1 shown in the drawing is an enclosed chamber that is rectangular in top plan and has a base wall 3, a pair of opposed end walls 5, a pair of opposed side walls 7, and a top cover 9.

15 The cell includes an inlet 11 for titania pellets in the top cover 9 near the left hand end of the cell as viewed in the drawing. This end of the cell is hereinafter referred to as "the rearward end" of the cell. The pellets are formed in a "green" state in a pin mixer 20 51 and are then sintered in a sintering furnace 53 and thereafter are stored in a storage bin 55. Pellets from the storage bin 55 are supplied via a vibratory feeder 57 to the cell inlet 11.

25 The cell further includes an outlet 13 for titanium metal pellets in the base wall 3 near the right hand end of the cell as viewed in the drawing. This end of the cell is hereinafter referred to as "the forward end" of the cell. The outlet 13 is in the form of a sump 30 defined by downwardly converging sides 15 and an upwardly inclined auger 35 arranged to receive titanium pellets from a lower end of the sump and to transport the pellets away from the cell.

35 The cell contains a bath 21 of molten electrolyte. The preferred electrolyte is CaCl₂ with at least some CaO.

- 14 -

The cell further includes an anode 23 in the form of a graphite block extending into the bath 21 and supported so that the block can be progressively lowered 5 into the bath 21 as lower sections of the anode graphite are consumed by cell reactions at the anode.

The cell further includes a cathode 25 in the form of a plate that is immersed in the bath 21 and is 10 positioned a short distance above the base wall 3. The cathode plate 25 is supported in the cell so that the upper surface of the cathode plate 25 is horizontal or slightly inclined downwardly from the rearward end to the forward end of the cell. The length dimension of the 15 cathode plate 25 is selected having regard to the residence time required for pellets in the bath. The width dimension of the cathode plate 25 is selected having regard to the total production required. The cathode plate 25 is supported to move in the forward and rearward 20 directions in an oscillating motion.

The applicant has found that movement of the cathode plate 25 in a repeated sequence that comprises a short period of oscillating motion and a short rest period 25 can cause pellets on the upper surface of the cathode plate 25 to move over the upper surface in a series of short steps from the rearward end to the forward end of the cell.

Moreover, the applicant has found that the above-described type of motion can cause pellets across the width of the cathode plate 25 to move at a constant rate so that the pellets have substantially the same residence time within the bath 21.

35

More particularly, the cell is arranged so that titania pellets supplied to the cell via the inlet 11 fall

- 15 -

downwardly onto the upper surface of the cathode plate 25 near the rearward end of the cell and are caused to move forwardly over the upper surface of the cathode plate 25 and fall off the forward end of the cathode plate 25 into 5 the outlet 13. More particularly, the cell is arranged so that, in use, the pellets move forwardly over the upper surface of the cathode plate 25 as a closely packed monolayer. In order to achieve close packing of the pellets, the cathode plate 25 includes an upstanding lip (not 10 shown) at the forward end thereof that causes pellets to build-up behind the lip along the length of the cathode plate 25.

15 The applicant has found that it is preferable that the titania pellets be substantially round since it is possible to cause these pellets to move over the upper surface of the cathode plate 25 in a more predictable manner than is possible with more angular pellets.

20 In addition, the applicant has found that it is undesirable that the pellets "stick" to the upper surface of the plate to an extent that inhibits forward movement of the pellets and that the pellets "stick" together. These considerations support the preference for round 25 pellets. It is relevant to note that oscillating movement of the cathode plate 25 minimises sticking of pellets. In addition, the plate may be coated with materials such as tantalum and titanium diboride to minimise sticking.

30 The applicant has also found that the size and weight of the pellets should be selected so that the pellets settle quite quickly onto the upper surface of the cathode plate 25 and do not become suspended in the electrolyte in the molten bath 21.

35

In overall terms, it is preferable to select the smallest possible pellet size that can move over the

- 16 -

cathode plate 25 in an efficient manner, i.e. without sticking to the plate, in order to optimise mass throughput of the cell.

5 The cell further includes a power source 31 for applying a potential across the anode block 23 and the cathode plate 25 and an electrical circuit that electrically interconnects the power source 31, the anode block 23, and the cathode plate 25. The electrical 10 circuit is arranged so that the power source 31 is connected to the rearward end of the cathode plate 25.

In use of the cell, titania pellets are supplied to the upper surface of the cathode plate 25 at the 15 rearward end of the cell so as to form a mono-layer of pellets on the cathode plate 25 and the plate is moved as described above and causes the pellets to step forward over the surface of the plate to the forward end of the cell and ultimately fall from the forward end of the 20 plate. The pellets are progressively electrochemically reduced in the cell as the pellets are moved over the surface of the cathode plate 25. The operating parameters of the cathode plate 25 are selected so that the pellets have sufficient residence time in the cell to achieve a 25 required level of reduction of the titania pellets. Typically, 2-4 mm titania pellets require 4 hours residence time to be reduced to titanium with a concentration of 0.3 wt% oxygen at a cell operating voltage of 3 V.

30 The applicant has found that the above-described arrangement results in substantial reduction of titania pellets within a short distance from the forward end of the cell.

35 The applicant has found that there are a number of factors that have an impact on the overall operation of

- 17 -

the cell. Some of these factors, namely pellet size and shape and motion of the cathode plate 25, are discussed above. Another relevant factor is the exposed surface areas of the upper surface of the cathode plate 25 and the 5 anode block 23. On the basis of work to date, the applicant believes that larger rather than smaller cathode plates 25 in relation to the exposed surface area of the anode block 23 is preferable. In other words, the applicant believes that a larger rather than a smaller 10 anodic current density is preferable.

In use of the cell, the anode block 23 is progressively consumed by a reaction between carbon in the anode block 23 and O⁻ anions generated at the cathode 15 plate 25, and the reaction occurs predominantly at the lower edges of the anode block 23.

It is preferred that the distance between the upper surface of the cathode plate 25 and the lower edges 20 of the anode block 23 be maintained substantially constant in order to minimise changes that may be required to other operating parameters of the process. Consequently, the cell further includes a means (not shown) for progressively lowering the anode block into the 25 electrolyte bath 21 to maintain the distance between the upper surface of the cathode plate 25 and the lower edges of the anode block 23 substantially constant.

Preferably the distance between the upper surface 30 of the cathode plate 25 and the lower edges of the anode block 23 is selected so that there is sufficient resistance heating generated to maintain the bath 21 at a required operating temperature.

35 Preferably the cell is operated at a potential that is above the decomposition potential of CaO. Depending on the circumstances, the potential may be as

- 18 -

high as 4-5V. In accordance with the above-described mechanism, operating above the decomposition potential of CaO facilitates deposition of Ca metal on the cathode plate 25 due to the presence of Ca⁺⁺ cations and migration 5 of O⁻⁻ anions to the anode block 23 as a consequence of the applied field and reaction of the O⁻⁻ anions with carbon of the anode block 23 to generate carbon monoxide and carbon dioxide and release electrons. In addition, in accordance with the above-described mechanism, the deposition of Ca 10 metal results in chemical reduction of titania via the mechanism described above and generates O⁻⁻ anions that migrate to the anode block 23 as a consequence of the applied field and further release of electrons. Operating the cell below the decomposition potential of CaCl₂, 15 minimises evolution of chlorine gas, and is an advantage on this basis.

As is indicated above, the operation of the cell generates carbon monoxide and carbon dioxide and 20 potentially chlorine-containing gases at the anode and it is important to remove these gases from the cell. The cell further includes an off-gas outlet 41 in the top cover 9 of the cell and a gas treatment unit 43 that treats the off-gases before releasing the treated gases to 25 atmosphere. The gas treatment includes removing carbon dioxide and any chlorine gases and may also include combusting carbon monoxide to generate heat for the process.

Titanium pellets, together with electrolyte that 30 is retained in the pores of the titanium pellets, are removed from the cell continuously or semi-continuously at the outlet 13. The discharged material is transported via the auger 35 to a water spray chamber 37 and quenched to a 35 temperature that is below the solidification temperature of the electrolyte, whereby the electrolyte blocks direct exposure of the metal and thereby restricts oxidation of

- 19 -

the metal. The discharged material is then washed to separate the retained electrolyte from the metal powder. The metal powder is thereafter processed as required to produce end products.

5

The above-described cell and process are an efficient and an effective means of continuously and semi-continuously electrochemically reducing metal oxides in the form of pellets to produce metal having a low oxygen
10 concentration

Specifically, the electrolytic cell shown in the drawing is one example only of a large number of possible cell configurations that are within the scope of the
15 present invention.

- 20 -

CLAIMS:

1. A process for electrochemically reducing metal oxide powders and/or pellets in an electrolytic cell that includes a bath of molten electrolyte, a cathode, and an anode, the cathode being in the form of a member having an upper surface for supporting metal oxide powders and/or pellets that is horizontally disposed or slightly inclined and has a forward end and a rearward end and is immersed in the electrolyte bath and is supported for movement so as to cause metal oxide powders and/or pellets on the upper surface of the cathode to move toward the forward end of the member, which process includes the steps of:
 - (a) applying a cell potential across the anode and the cathode that is capable of electrochemically reducing metal oxide supplied to the molten electrolyte bath,
 - (b) continuously or semi-continuously feeding metal oxide powders and/or pellets into the molten electrolyte bath so that the powders and/or pellets deposit on an upper surface of the cathode,
 - (c) causing metal oxide powders and/or pellets to move over the upper surface of the cathode toward the forward end of the cathode while in contact with molten electrolyte whereby electrochemical reduction of the metal oxide to metal occurs as the powders and/or pellets move toward the forward end, and
 - (d) continuously or semi-continuously removing at least partially electrochemically reduced metal oxide powders and/or pellets from the molten electrolyte bath.
- 30 2. The process defined in claim 1 wherein step (b) includes feeding the metal oxide powders and/or pellets into the molten electrolyte bath so that the powders and/or pellets form a layer that is one or two particles deep on the upper surface of the cathode.
- 35 3. The process defined in claim 1 wherein step (b) includes feeding the metal oxide powders and/or pellets

- 21 -

into the molten electrolyte bath so that the powders
and/or pellets deposit as a pile of powders and/or pellets
on the upper surface of the cathode and step(c) causes the
powders and/or pellets in the pile to be shaken out into a
5 layer that that is one or two particles deep and moves
over the upper surface of the cathode toward the forward
end of the cathode.

4. The process defined in claim 1 wherein step (c)
10 includes causing metal oxide powders and/or pellets to
move on the upper surface of the cathode toward the
forward end of the cathode as a layer of powders and/or
pellets that is one or two particles deep.

15 5. The process defined in any one of the preceding
claims wherein step (c) includes selectively moving the
cathode so as to cause metal oxide powders and/or pellets
on the upper surface of the cathode to move toward the
forward end of the cathode.

20 6. The process defined in claim 5 wherein step (c)
includes moving the cathode in forward and rearward
directions so as to cause metal oxide powders and/or
pellets on the upper surface of the cathode to move toward
25 the forward end of the cathode.

7. The process defined in claim 6 includes moving
the cathode in a repeated sequence that comprises a short
period of oscillating motion in the forward and rearward
30 directions and a short rest period.

8. The process defined in any one of the preceding
claims wherein step (c) includes moving the cathode so as
to cause powders and/or pellets across the width of the
35 cathode to move at the same rate so that the powders
and/or pellets have substantially the same residence time
within the bath.

9. The process defined in any one of the preceding claims includes washing powders and/or pellets that are removed from the cell and separating electrolyte that is 5 carried from the cell with the pellets.

10. 10. The process defined in claim 9 includes recovering electrolyte that is washed from the powders and/or pellets and recycling the electrolyte to the cell.

11. 15. The process defined in any one of the preceding claims includes applying a cell potential above a decomposition potential of at least one constituent of the electrolyte so that there are cations of a metal other than that of the cathode metal oxide in the electrolyte.

12. 20. The process defined in claim 11 wherein, in a situation in which the metal oxide is titania the electrolyte be a CaCl₂-based electrolyte that includes CaO as one of the constituents, the process includes maintaining the cell potential above the decomposition potential for CaO.

13. 25. The process defined in any one of the preceding claims wherein the particle size of the powders and/or pellets is in the range of 0.5-4 mm.

14. 30. An electrolytic cell for electrochemically reducing metal oxide powders and/or pellets, which electrolytic cell includes (a) a bath of a molten electrolyte, (b) a cathode in the form of a member having an upper surface for supporting metal oxide powders and/or pellets that is horizontally disposed or slightly inclined and has a forward end and a rearward end and is immersed 35 in the electrolyte bath and is supported for movement so as to cause metal oxide powders and/or pellets on the upper surface of the cathode to move toward the forward

- 23 -

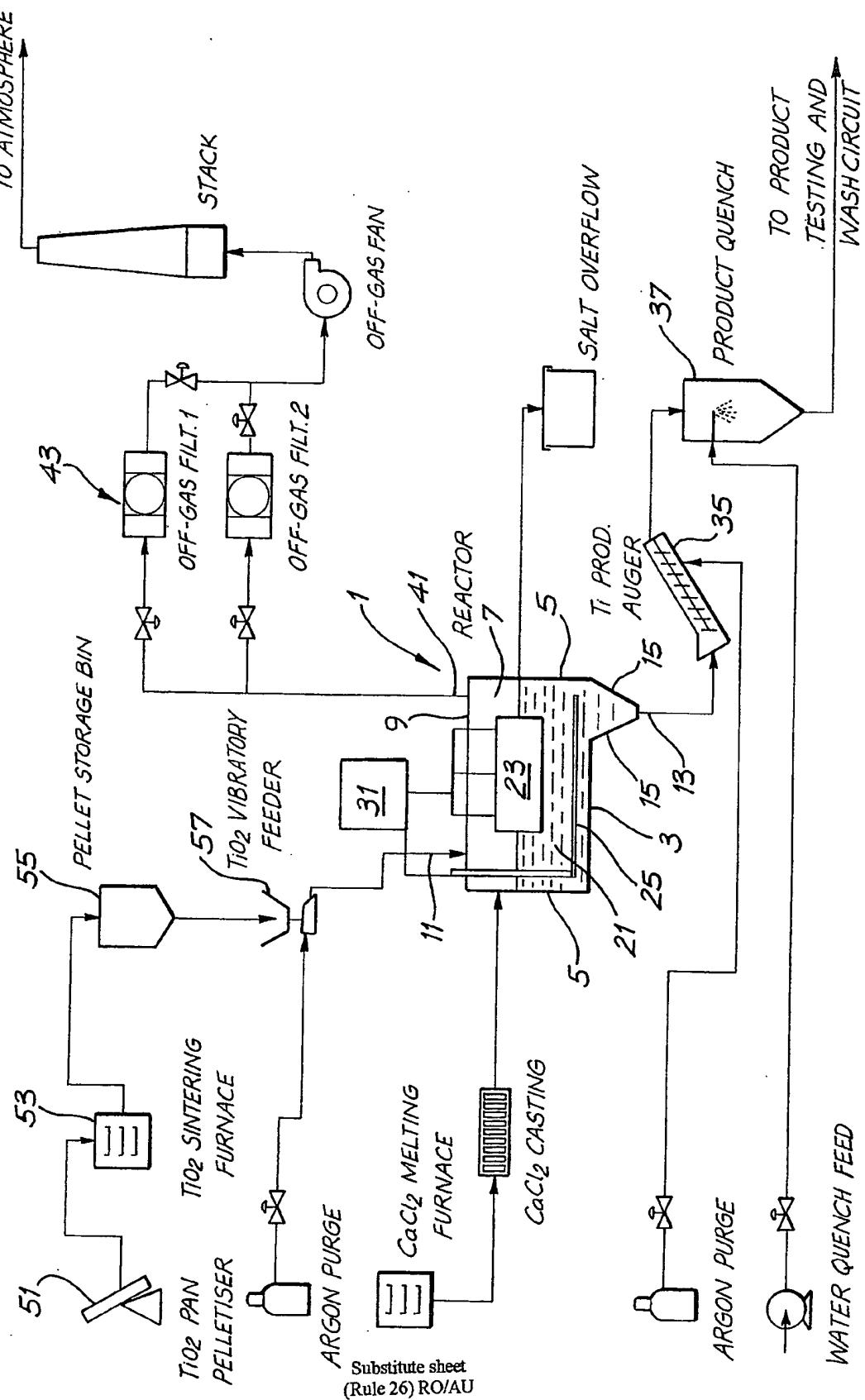
- end of the cathode, (c) an anode, (d) a means for applying a potential across the anode and the cathode, (e) a means for supplying metal oxide powders and/or pellets to the electrolyte bath so that the metal oxide powders and/or 5 pellets can deposit onto an upper surface of the cathode, (f) a means for causing metal oxide powders and/or pellets to move over the upper surface of the cathode toward the forward end of the cathode while in contact with molten electrolyte whereby electrochemical reduction of the metal 10 oxide to metal can occur as the powders and/or pellets move toward the forward end, and (g) a means for removing at least partially electrochemically reduced metal oxides from the electrolyte bath.
- 15 15. The cell defined in claim 14 wherein the cathode is a plate.
16. The cell defined in claim 14 or claim 15 wherein the means for causing metal oxide powders and/or pellets 20 to move over the upper surface of the cathode includes a means for moving the cathode so as to cause movement of metal oxide powders and/or pellets.
17. The cell defined in claim 16 wherein the means for causing metal oxide powders and/or pellets to move over the upper surface of the cathode includes a means for moving the cathode in forward and rearward directions. 25
18. The cell defined in any one of claims 14 to 17 wherein the cathode is formed to cause metal oxide powders and/or pellets to move on the upper surface of the cathode toward the forward end of the cathode as a layer of powders and/or pellets that is one or two particles deep. 30
- 35 19. The cell defined in claim 18 wherein the cathode is formed with an upstanding lip at the forward end that causes powders and/or pellets to build-up behind the lip.

20. The cell defined in claim 18 or claim 19 wherein
the upper surface of the cathode is formed with a series
of transversely extending grooves that promote close
5 packing of the powders and/or pellets.

21. The cell defined in any one of claims 14 to 20
wherein the means for applying an electrical potential
across the anode and the cathode includes an electrical
10 circuit in which a power source is connected to a forward
end of the cathode.

22. The cell defined in any one of claims 14 to 21
wherein the anode extends downwardly into the electrolyte
15 bath and is positioned a predetermined distance above the
upper surface of the cathode.

23. The cell defined in claim 22 includes a means for
moving the anode downwardly into the electrolyte bath as
20 the anode is consumed to maintain the predetermined
distance between the anode and the cathode.



INTERNATIONAL SEARCH REPORT

International application No.
PCT/AU2004/000809

A. CLASSIFICATION OF SUBJECT MATTER		
Int. Cl. ⁷ : C25C 5/00, C22B 9/14, 34/12		
According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED		
Minimum documentation searched (classification system followed by classification symbols) IPC ⁷ as above		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) Derwent WPI: IPC ⁷ as above and electrol+		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
E, X	WO 2004/053201 A1 (BHP Billiton Innovations Pty Ltd) 24 June 2004 Whole Document	1 to 23
P, A	Derwent Abstract Accession No. 2004-185614/18, Class M25 (M28), JP 2004-052037 A (Nippon Light Metal Co) 19 February 2004 Abstract and Diagrams	
A	Derwent Abstract Accession No. 2003-372523/35, Class M25, WO 2003/038156 A1 (Nippon Light Metal Co) 8 May 2003 Abstract and Diagrams	
A	US 6540902 B1 (Redey et al) 1 April 2003 Whole Document	
<input checked="" type="checkbox"/> Further documents are listed in the continuation of Box C <input checked="" type="checkbox"/> See patent family annex		
<small>* Special categories of cited documents:</small> "A" document defining the general state of the art which is not considered to be of particular relevance "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "E" earlier application or patent but published on or after the international filing date "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "O" document referring to an oral disclosure, use, exhibition or other means "&" document member of the same patent family "P" document published prior to the international filing date but later than the priority date claimed		
Date of the actual completion of the international search 14 July 2004	Date of mailing of the international search report 20 JUL 2004	
Name and mailing address of the ISA/AU AUSTRALIAN PATENT OFFICE PO BOX 200, WODEN ACT 2606, AUSTRALIA E-mail address: pct@ipaaustralia.gov.au Facsimile No. (02) 6285 3929	Authorized officer DAVID K. BELL Telephone No : (02) 6283 2309	

INTERNATIONAL SEARCH REPORT

International application No. PCT/AU2004/000809
--

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P, A	US 2004/0060826 A1 (Godfrey) 1 April 2004 Whole Document	

INTERNATIONAL SEARCH REPORT

Information on patent family members

International application No.

PCT/AU2004/000809

This Annex lists the known "A" publication level patent family members relating to the patent documents cited in the above-mentioned international search report. The Australian Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

Patent Document Cited in Search Report	Patent Family Member
WO 2004053201	
JP 2004052037	
WO 03038156	JP 2003129268
US 6540902	
US 2004060826	WO 2004029309

Due to data integration issues this family listing may not include 10 digit Australian applications filed since May 2001.

END OF ANNEX